Atomic Structure Notes

Electrons – \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ charged part of atom found in electron clouds

Nucleus – \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_charged center

* + Protons – positively charged particles in the nucleus
  + Neutrons – particles in the nucleus that have \_\_\_\_\_\_ charge; neutral
  + Charge of protons and electrons are opposite but equal, so their charges cancel each other out.
* Atoms of different elements contain different numbers of \_\_\_\_\_\_\_\_\_\_\_\_

Atomic Number – number of protons in an atom

Ex: Carbon – atomic number is 6

\_\_\_\_\_\_ protons, \_\_\_\_\_\_ electrons

Atomic Mass – Total number of protons and neutrons in an atom.

Ex: Carbon – mass number is 12, atomic number is 6.

12 – 6 = \_\_\_\_\_

There are six neutrons.

Let’s draw an atom of Boron–

Atomic Number is 5

Atomic Mass is 11 (10.81 rounds to 11)

**Nucleus – protons and neutrons**

**electrons**

Protons = \_\_\_\_\_\_

Electrons = \_\_\_\_\_

Neutrons = 11-5 = \_\_\_\_