Atomic Structure Notes

Electrons – \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ charged part of atom found in electron clouds

Nucleus – \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_charged center

* + Protons – positively charged particles in the nucleus
	+ Neutrons – particles in the nucleus that have \_\_\_\_\_\_ charge; neutral
	+ Charge of protons and electrons are opposite but equal, so their charges cancel each other out.
* Atoms of different elements contain different numbers of \_\_\_\_\_\_\_\_\_\_\_\_

Atomic Number – number of protons in an atom

 Ex: Carbon – atomic number is 6

 \_\_\_\_\_\_ protons, \_\_\_\_\_\_ electrons

Atomic Mass – Total number of protons and neutrons in an atom.

 Ex: Carbon – mass number is 12, atomic number is 6.

 12 – 6 = \_\_\_\_\_

 There are six neutrons.

Let’s draw an atom of Boron–

 Atomic Number is 5

 Atomic Mass is 11 (10.81 rounds to 11)

**Nucleus – protons and neutrons**

**electrons**

Protons = \_\_\_\_\_\_

Electrons = \_\_\_\_\_

Neutrons = 11-5 = \_\_\_\_